

The Crystal and Molecular Structure of Tetra- μ -chloro-tetra- μ -ethyloxycarbene-di- μ_3 -hydroxo dodecacarbonyl-di-chloro-hexaruthenium(II)

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The crystal and molecular structure of tetra- μ -chloro-tetra- μ -ethyloxycarbene-di- μ_3 -hydroxo-dodecacarbonyl-di-chloro-hexaruthenium(II)-benzene, $[\text{Ru}_3\text{Cl}_3(\text{COC}_2\text{H}_5)_2(\text{OH})(\text{CO})_6]_2 \cdot \text{C}_6\text{H}_6$, have been determined. The complex crystallizes as monoclinic crystals in space group $P 2_1/c$ with unit cell dimensions $a = 10.71(2)$ Å, $b = 10.41(2)$ Å, $c = 21.92(3)$ Å, $\beta = 113^\circ 15' (10')$ and $Z = 2$. A three-dimensional X-ray analysis, based on 3101 observed reflections, converged to a R of 0.076. The crystal structure of the complex consists of centrosymmetric hexanuclear molecules. The three ruthenium atoms of the asymmetric unit are octahedrally coordinated. The octahedra share a common vertex occupied by a hydroxo group which bridges three ruthenium atoms. Significant differences in the Ru–Cl bond lengths are accounted by the “trans influence” of the carbene nature of the $\text{C}_2\text{H}_5\text{CO}$ groups.

Introduction

η -Allyl ruthenium complexes of formula $\text{RuX}(\text{C}_3\text{H}_5)(\text{CO})_3$ ($\text{X} = \text{Cl}, \text{Br}$) have been used as catalysts in several reactions involving unsaturated substrates [1, 2]. The reactivity with different olefinic and acetylenic ligands was also widely investigated [3].

In the course of these studies some new ruthenium complexes, obtained in crystalline and amorphous forms, by reaction of $\text{RuCl}(\text{C}_3\text{H}_5)(\text{CO})_3$ (I) with ethylene were isolated [2].

Since the analytical and spectroscopic data (IR and ^1H n.m.r.) did not allow to establish exactly the nature of the products and to understand some IR and analytical differences observed for the crystalline and amorphous product, recourse was had to X-ray crystallography.

Experimental

Preparation

By reaction of I (1 g, 3.8 mmol) with ethylene under pressure (50–60 atm) at 70–80 °C for 10 h in benzene solution and by standing the reaction mixture in the autoclave at room temperature for several days, a crystalline compound (II), initially formulated as $[\text{RuCl}(\text{COC}_2\text{H}_5)(\text{CO})_2]_n$ with solvate benzene, was obtained [Found: C, 24.57; H, 2.17; Cl, 14.3; Ru, 40.2; M(benzene, Mechrolab Osmometer), 1600. Calc. for $\text{C}_{30}\text{H}_{28}\text{Cl}_6\text{O}_{18}\text{Ru}_6$ (on the basis of the formula determined by X-ray structural data): C, 24.09; H, 1.88; Cl, 14.22; Ru, 40.55%; M, 1495.7]. The complex could also be isolated as an amorphous product (III) either by adding *n*-pentane to the benzene solution or by evaporating to dryness [Found: C, 23.82; H, 2.18; Cl, 14.6; Ru, 39.2. $\text{C}_5\text{H}_5\text{ClO}_3\text{Ru}$ requires C, 24.05; H, 2.02; Cl, 14.21; Ru, 40.4%].

IR and ^1H n.m.r. Spectra

The IR spectra (Perkin–Elmer 225 spectrometer, in KBr) of the crystalline and amorphous products showed three bands in the carbonyl stretching region at 2130 vw, 2060vs, 2005vs cm^{-1} and the bands assignable to a bridging acyl group (1543s cm^{-1}) [4] and to the terminal (329w cm^{-1}) and bridging chlorine atoms (290w and 260w cm^{-1}) [5]. The IR spectrum of the crystalline product differed from that of the amorphous one for a weak but sharp absorption at 3280 cm^{-1} [$\nu(\text{OH})$], by a more complex pattern of bands in the Ru–C–O deformation and Ru–C stretching region (650–400 cm^{-1}) and by the bands due to solvate benzene.

The ^1H n.m.r. spectrum of the crystalline product (Varian T 60 spectrometer, in benzene solution) showed a quartet and a triplet at 7.32 and 9.15 τ

TABLE I. Positional and Thermal Parameters for Non-hydrogen Atoms in $[\text{Ru}_3\text{Cl}_3(\text{COC}_2\text{H}_5)_2(\text{OH})(\text{CO})_6]_2 \cdot \text{C}_6\text{H}_6$ a,b.

	x	y	z	β_{11}	β_{22}	β_{33}	β_{12}	β_{13}	β_{23}
Ru(1)	-0.08384(10)	-0.20176(11)	0.40900(5)	81.7(10)	23.3(15)	20.6(2)	-5.1(8)	17.6(4)	-3.4(4)
Ru(2)	0.06520(10)	0.02238(12)	0.34289(5)	75.0(9)	40.2(15)	19.1(2)	0.8(8)	17.1(4)	0.0(4)
Ru(3)	-0.19243(10)	0.15187(12)	0.37316(5)	70.4(9)	36.0(15)	20.4(2)	1.9(8)	14.1(4)	-1.1(4)
Cl(1)	0.13394(34)	-0.20980(38)	0.50463(16)	88.4(32)	44.7(51)	20.5(8)	9.8(28)	16.0(13)	-1.6(14)
Cl(2)	0.03363(37)	-0.21990(39)	0.33623(18)	108.0(36)	33.5(50)	25.1(9)	3.4(30)	28.4(15)	-5.8(15)
Cl(3)	-0.19267(36)	-0.15407(42)	0.48031(17)	99.0(35)	58.1(53)	23.3(8)	-4.0(32)	25.7(14)	-1.6(15)
C(1)	-0.1038(16)	-0.3769(18)	0.4170(8)	107(16)	75(26)	24(4)	-38(15)	16(6)	1(7)
C(2)	-0.2436(15)	-0.2023(16)	0.3344(8)	93(14)	47(21)	26(4)	-23(12)	27(6)	-10(6)
C(3)	0.1465(15)	0.0427(18)	0.2829(8)	90(15)	92(25)	23(4)	9(14)	14(6)	-3(7)
C(4)	0.2430(14)	0.0003(17)	0.4077(7)	84(14)	83(23)	24(4)	-7(13)	22(6)	-4(7)
C(5)	-0.3455(17)	0.0640(17)	0.3741(8)	122(19)	41(22)	30(4)	11(15)	19(7)	-6(7)
C(6)	-0.3053(15)	0.2876(17)	0.3438(7)	99(15)	93(25)	20(3)	1(14)	21(6)	6(6)
C(7)	-0.2311(14)	0.0923(16)	0.2809(6)	92(13)	56(20)	14(3)	-8(12)	9(5)	1(5)
C(8)	0.0700(11)	0.2065(15)	0.3584(6)	52(11)	67(22)	23(3)	19(10)	14(5)	6(6)
C(9)	-0.3624(15)	0.1058(19)	0.2217(8)	77(14)	108(27)	22(4)	0(14)	-1(6)	7(7)
C(10)	-0.3629(18)	0.2253(19)	0.1834(9)	136(20)	65(25)	28(5)	17(17)	20(8)	15(8)
C(11)	0.1766(15)	0.3009(18)	0.3580(8)	83(14)	101(26)	23(4)	-15(14)	17(6)	-8(7)
C(12)	0.1649(21)	0.4437(20)	0.3825(11)	166(26)	24(24)	51(8)	15(18)	31(12)	-4(9)
C(13)	0.4235(26)	0.0358(32)	0.0363(13)	171(32)	291(55)	36(7)	73(34)	14(12)	1(15)
C(14)	0.4899(34)	0.1302(29)	0.0120(13)	286(48)	165(45)	39(8)	-18(36)	14(16)	-19(14)
C(15)	0.4390(32)	-0.0909(34)	0.0246(14)	254(46)	243(53)	40(9)	-73(39)	4(16)	13(16)
O(1)	-0.1189(14)	-0.4803(14)	0.4243(7)	157(17)	86(19)	38(4)	-32(13)	29(7)	7(7)
O(2)	-0.3461(12)	-0.2025(13)	0.2887(6)	102(13)	127(20)	27(3)	-17(12)	-1(5)	-8(6)
O(3)	0.1997(12)	0.0459(14)	0.2474(6)	135(15)	148(20)	30(3)	18(13)	46(6)	9(6)
O(4)	0.3509(11)	-0.0135(13)	0.4438(6)	94(12)	130(20)	30(3)	25(12)	13(5)	8(6)
O(5)	-0.4450(11)	0.0221(13)	0.3705(7)	87(12)	107(19)	48(5)	-8(11)	35(6)	10(7)
O(6)	-0.3744(13)	0.3776(12)	0.3265(6)	151(16)	57(17)	32(3)	48(12)	17(6)	7(5)
O(7)	-0.1351(9)	0.0436(10)	0.2705(4)	82(9)	54(14)	14(2)	4(8)	8(3)	4(4)
O(8)	-0.0279(9)	0.2556(10)	0.3698(5)	78(9)	12(12)	28(3)	-8(8)	22(4)	1(4)
O(9)	-0.0431(9)	0.0020(10)	0.4069(4)	84(9)	44(13)	17(2)	7(8)	21(4)	4(4)

^a The numbers in parentheses in this and other tables are the estimated deviations in the least significant digits. The anisotropic thermal parameters are of the form $\exp[-10^{-4}(h^2\beta_{11} + k^2\beta_{22} + l^2\beta_{33} + 2hk\beta_{12} + 2hl\beta_{13} + 2kl\beta_{23})]$.

from TMS, in the proper intensity ratio for an ethyl group, and a resonance at 10.9 τ assignable to a triple bridging OH group [6]. This last resonance was absent in the spectrum of the amorphous product.

X-ray Crystallography

Unit-cell dimensions and space group for the crystals of the complex II were obtained from precession photographs taken with Mo K α radiation. The crystals belonged to the monoclinic system with $a = 10.71(2)$ Å, $b = 10.41(2)$ Å, $c = 21.92(3)$ Å, $\beta = 113^\circ 15'$ ($10'$). Systematic absences ($h0l$ absent for $l = 2n + 1$ and $0k0$ absent for $k = 2n + 1$) indicated space group P $2_1/c$. For $Z = 2$ the calculated density was 2.45 g/cm³.

The intensity data were recorded with Ni-filtered Cu K α radiation ($\lambda = 1.5418$ Å) by means of Weissenberg photographs with multiple film technique and integration process. A crystal elongated in the b direction was reduced to nearly cylindrical shape with a radius of 0.013 cm ($\mu R = 2.70$ for Cu K α radiation). Nine layers with b as rotation axis ($k = 0$ through 8) were taken.

A total of 4138 independent reflections (3101 observed) were collected. The intensities measured with a Nonius microdensitometer have been corrected for Lorentz and polarization effects, for the incipient but incomplete $\alpha_1 - \alpha_2$ spot doubling and for the absorption factor using the three-constant formula proposed by Palm [7] for cylindrical crystals.

Solution and Refinement of the Structure

The ruthenium atoms were located from three-dimensional Patterson function. All chlorine, carbon, and oxygen atoms were located by successive Fourier syntheses. Various cycles of block-matrix least squares refinement lowered the value of $R = \Sigma ||F_o| - |F_c|| / \Sigma |F_o|$ to 0.092: unit weights were assumed for all observed reflections, isotropic temperature factors were attributed to carbon and oxygen atoms, while anisotropic temperature factors were attributed to ruthenium and chlorine atoms.

The anisotropic refinement was extended to all the atoms and the weighting scheme $\sqrt{w} = 1/(a + |F_o| + b|F_o|^2)$ proposed by Cruickshank [8] was introduced with $a = 40.0$ and $b = 0.005$.

Seventeen strong reflections, probably affected by extinction, were excluded from the last refinement cycles. The final R value was 0.076.

Scattering factors used in the structure factor calculations were taken from International Tables for X-ray Crystallography [9]. All the programs used are incorporated in the X-ray 70 system [10]. Positional and thermal parameters for the atoms in the asymmetric unit are given in Table I. Observed and calculated structure factors are compared in Table II.

Results and Discussion

The resolution of the structure of the crystalline product II allows to formulate this as tetra- μ -chloro-tetra- μ -ethyloxycarbene-di- μ_3 -hydroxo-dodecacarbonyl-dichloro-hexaruthenium(II)-benzene, $[\text{Ru}_3\text{Cl}_3(\text{COC}_2\text{H}_5)_2(\text{OH})(\text{CO})_6]_2 \cdot \text{C}_6\text{H}_6$.

The crystal structure of II is composed of two discrete hexanuclear molecules located at the inversion centers at (0, 0, 1/2) and (0, 1/2, 0) in the monoclinic unit cell. At another pair of centers at (1/2, 0, 0) and (1/2, 1/2, 1/2) are two molecules of benzene of crystallization.

This structure is closely analogous to that of a toluene-containing form of di- μ -aquo-di- μ_3 -hydroxo-deca(1,1,1-trifluoropentane-2,4-dionato)hexanickel(II) [11] which also crystallizes in the space group P $2_1/c$ with two hexanuclear molecules at two inversion centers and two disordered molecules of toluene at another pair of centers. The resemblance goes further as both the complex molecules are composed of two trinuclear units, related by a center of symmetry and joined by two bridging water molecules in the nickel complex and by two bridging chlorine atoms in II. In both cases the three metal atoms in the asymmetric unit are each octahedrally coordinated and the three octahedra share a common vertex at the position of the hydroxo group.

Fig. 1 is an ORTEP [12] drawing of the molecule. Bond lengths and angles are given in Tables III and IV, together with their standard deviations.

As shown the molecular structure of II has a number of unusual features that seldom are present in a same molecule.

The first one is the presence of two bridging $\text{C}_2\text{H}_5\text{-CO}$ groups between the Ru(2) and Ru(3) atoms; a similar situation was up to now found only in the crystal structure of bis-(μ -phenyloxycarbene-tricarbonyliron) [13]. In that case, however, the C atoms of the two bridging groups are both linked to the same iron atom, whereas the oxygen atoms are linked to the other iron atom. The C-O bond lengths of $\text{C}_2\text{H}_5\text{CO}$ groups [C(7)-O(7): 1.246 Å; C(8)-O(8): 1.276 Å] longer than that of a keto group (1.23 Å) and analogous to those found in the "carbene compound" $\text{Fe}_2(\text{CO})_6[\text{C}(\text{O})\text{Ph}]_2$ [13] and the Ru-C distances [Ru(2)-C(8): 1.944 Å; Ru(3)-C(7): 1.996 Å], longer than the ruthenium carbonyl lengths (1.89 Å) but shorter than those of a Ru-C (σ) bond (2.07 Å) [14], indicate for the bridging $\text{C}_2\text{H}_5\text{CO}$ groups an arrangement of carbene type [13]. The carbene character of $\text{C}_2\text{H}_5\text{C}(\text{O})$ groups, which display a high σ -donor and a low π -acceptor power, is also evidenced by the lengthening by *trans* effect [15, 16] of the Ru(2)-Cl(2) and Ru(3)-Cl(1)ⁱ bonds.

The C_2H_5 groups in II arise from the insertion of a C_2H_4 molecule into a ruthenium-hydrogen bond formed by β -elimination of pentadienes from the

1, 7, 7, L												7, 7, L												5 936 891 -12 566 -552 -16 323 -398 10 830 -738 -11 951 -923 -9 455 -397																							
0 562 -649												12556 -1228												6 587 -576 -17 202 235 -19 170 -209 11 664 570 -14 283 -302 -9 599 -610																							
1 340 331												-514 -465												8 481 -520 -20 196 180 -19 352 490 14 625 620 -16 232 193 -10 469 461																							
2 505 495												2 225 -120 10 277 298												9 625 -664 -22 194 200																							
3 540 475												7 565 -1584												14 178 -192												P, 7, L											
4 941 -903												4 989 908												15 308 -354												11, 7, L											
5 772 -727												5 338 -281												16 135 219												0 225 194 -3 233 -271 -2 694 -640 -22 132 183											
6 248 -170												7 826 800 -1 1014 -952												1 733 798 -4 168 -205 -3 222 198												5, 8, L											
7 890 -374												8 243 214 -3 393 315												2 38P -422 -6 240 247 -5 1486 1405												0 643 564											
8 1210 1218												9 771 352 -5 887 811												4 30P 338 -7 284 -367 -6 656 568												2 667 548											
9 1066 1066												11 444 -433 -7 351 -295												5 267 180 -9 238 -268 -7 1410 1314												2 667 548											
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13 230 -210												14 270 260 -10 443 385												9 174 -195 -12 230 262 -13 1228 1268												5 644 613											
15 620 -578												16 746 389 -11 354 -316												-1 261 221 -14 339 441 -14 500 -526												6 501 470											
16 359 369												18 753 174 -12 705 -679												-2 555 580 -15 351 -440 -15 380 461												7 538 526											
17 243 254												-1 1338 1312 -13 333 390 -3 737 729												-16 197 -286 -17 462 -502												9 790 820											
19 269 245												-2 335 300 -16 370 388												-4 470 -397 -17 153 -312 -18 404 485												11 475 544											
-1 355 412												-3 797 749 -19 348 358												-6 626 -656												-19 366 485											
-2 370 -209												-4 1172 1103 -22 195 229												-7 1037 1068												-20 155 120											
-3 399 391												-5 667 -630 -23 395 -489												-8 555 496												-22 184 -281 -23 322 -165											
-4 848 917												-6 595 -590 -24 188 -299												-9 673 685												1 1306 -1903											
-5 212 170												-9 267 -219												-11 790 -179												2 241 -262											
-7 594 -595												-10 730 160												-12 404 -383												3 1157 -1192											
-8 720 -676												-11 676 374												-14 264 -296												4 325 -215											
-9 880 -917												-11 666 -619												0 513 -454												5 1539 1435											
-11 312 236												-12 593 -571												1 333 -320												-18 150 -171											
-13 902 862												-14 501 -579												2 336 318												-19 270 247											
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-16 665 666												-16 641 589												4 323 216												-21 202 -239											
-17 660 -639												-17 229 200												5 580 573												11 738 636											
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alkyl intermediate derived by a primary insertion of ethylene into a ruthenium-allyl bond of I [2].

The second feature is the presence of a triple bridging hydroxo group: such a group has been previously characterized in the crystal structure of [Pt(OH)(CH₃)₃]₄ [17, 18], Ni₆(CF₃COCHCOCH₃)₁₀(OH)₂(H₂O)₂ [11] and [W(OH)(CO)₃H]₄·4 OPPh₂-Et [19]. As the hydroxo group is a five-electron donor, in one molecule of II there are a total of sixty ligand-donated electrons, so each ruthenium atom can conform to the closed shell configuration without any metal-metal bond.

The close approach O(9)...Cl(3)^δ [3.174(9) Å] between the oxygen of the hydroxo group and the terminal chlorine atom is indicative of hydrogen bonding: as appears from Fig. 1, the relative orientation of the two atoms is favourable for hydrogen bonding which assures a further linking between the

two centrosymmetrically related trinuclear units in the molecule. This is also confirmed by the sharp IR absorption at 3280 cm⁻¹, consistent with the O-H stretching frequency for an O...Cl bond distance of 3.19 Å [20]. The origin of the hydroxo group of II, indicated also by a resonance at 10.9 τ in the ¹H n.m.r. spectrum [6], is not clear: it might arise from the moisture of the commercial ethylene used as well as from decomposition of carbon monoxide [21].

The last feature of the molecular structure of II is the different coordination of the three ruthenium atoms in the asymmetric unit: in fact, Ru(2) and Ru(3) are identically coordinated, each being linked to two carbonyl groups, to the "carbene atom" of one ethyloxycarbene group and to the oxygen atom of the other ethyloxycarbene group, to a bridging chlorine atom and to the triple bridging hydroxo group. Ru(1) on the contrary is linked to three chlorine

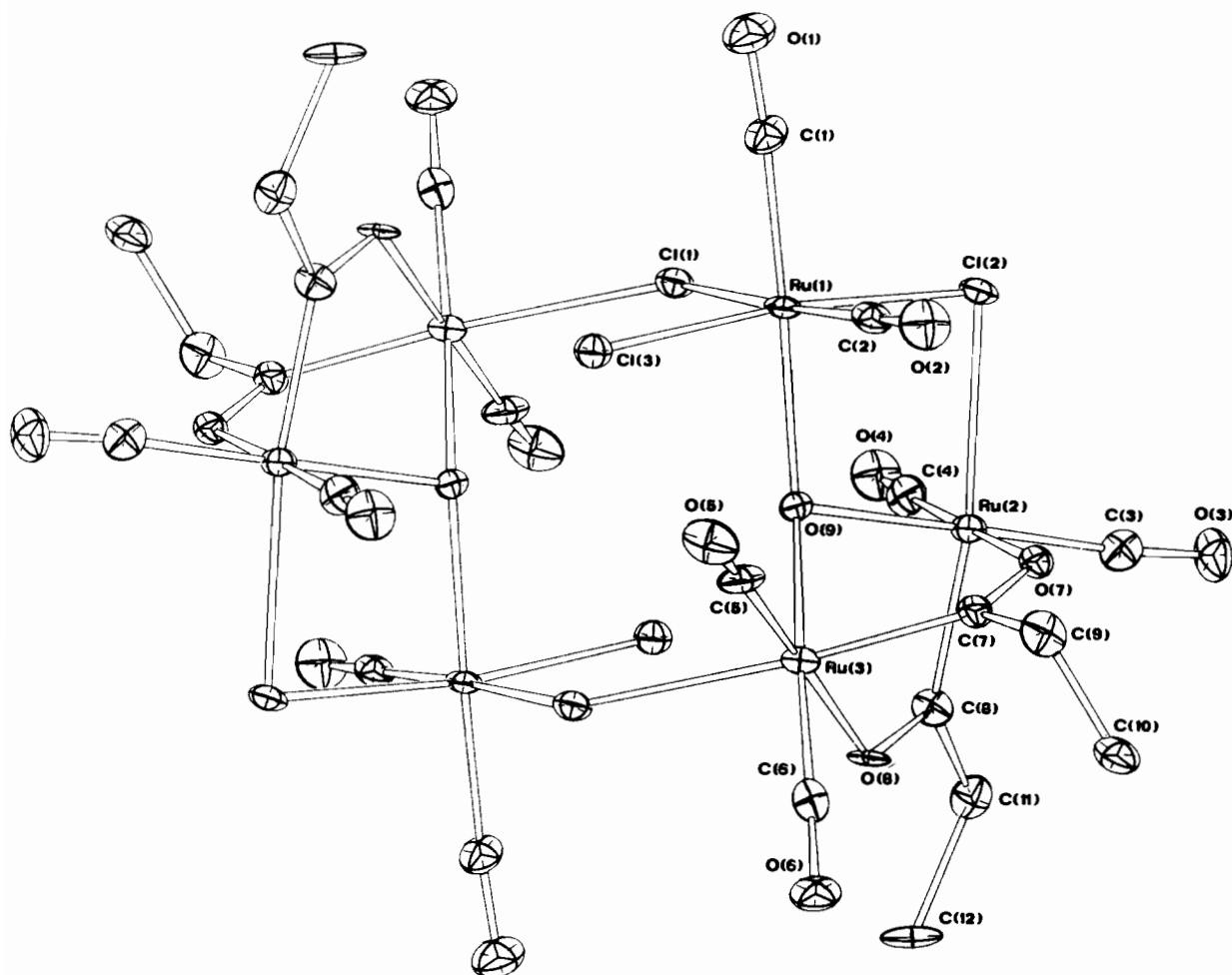


Figure 1. The molecular structure of $[\text{Ru}_3\text{Cl}_3(\text{COC}_2\text{H}_5)_2(\text{OH})(\text{CO})_6]_2 \cdot \text{C}_6\text{H}_6$.

TABLE III. Bond Distances (Å)^{a,b}.

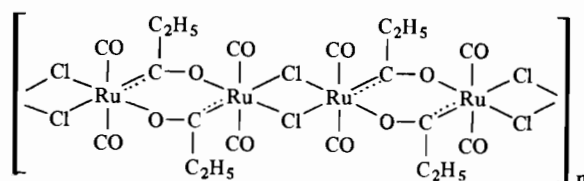
Ru(1)—Cl(1)	2.446(3)	Ru(3)—O(8)	2.094(10)
Ru(1)—Cl(2)	2.398(5)	Ru(3)—O(9)	2.146(9)
Ru(1)—Cl(3)	2.343(5)		
Ru(1)—C(1)	1.851(19)	C(1)—O(1)	1.109(24)
Ru(1)—C(2)	1.842(12)	C(2)—O(2)	1.158(16)
Ru(1)—O(9)	2.169(10)	C(3)—O(3)	1.134(25)
		C(4)—O(4)	1.123(16)
Ru(2)—Cl(2)	2.541(4)	C(5)—O(5)	1.125(23)
Ru(2)—C(3)	1.854(20)	C(6)—O(6)	1.162(21)
Ru(2)—C(4)	1.887(13)	C(7)—O(7)	1.246(19)
Ru(2)—C(8)	1.944(16)	C(8)—O(8)	1.276(19)
Ru(2)—O(7)	2.118(8)	C(7)—C(9)	1.497(17)
Ru(2)—O(9)	2.155(11)	C(9)—C(10)	1.499(28)
		C(8)—C(11)	1.509(23)
Ru(3)—Cl(1) ⁱ	2.571(4)	C(11)—C(12)	1.602(29)
Ru(3)—C(5)	1.885(19)	C(13)—C(14)	1.43(5)
Ru(3)—C(6)	1.805(17)	C(13)—C(15)	1.37(5)
Ru(3)—C(7)	1.996(14)	C(14)—C(15) ⁱⁱ	1.37(5)

^aThe figures in parentheses are the estimated standard deviations. ^bThe symmetry code used in this and in the subsequent table is the following:

i atom at $-x \ -y \ 1-z$
 ii atom at $1-x \ -y \ -z$

atoms (one terminal and two bridging), to two carbonyl groups and to the triple bridging hydroxo group.

Concerning the structure of the amorphous product, the IR spectrum does not show the presence of the triple bridging OH group and indicates a simpler arrangement of the CO groups as results from the lower Ru—C stretchings between 650 and 400 cm^{-1} . The high value of the molecular weight (~ 2000) furthermore suggests for this product a “ μ -catena” arrangement of $\text{RuCl}(\text{COC}_2\text{H}_5)(\text{CO})_2$ units of type:



Both the crystalline and amorphous products are stable in the solid state and give by treatment with

TABLE IV. Bond Angles ($^{\circ}$)^a.

Bond Angles at the Ruthenium Atoms			
Cl(1)–Ru(1)–Cl(2)	89.7(1)	O(9)–Ru(2)–C(8)	88.1(6)
Cl(1)–Ru(1)–Cl(3)	89.6(1)	C(4)–Ru(2)–O(7)	178.9(11)
Cl(1)–Ru(1)–C(1)	89.5(4)	C(4)–Ru(2)–C(3)	86.0(7)
Cl(1)–Ru(1)–C(2)	176.6(5)	C(4)–Ru(2)–C(8)	92.2(6)
Cl(1)–Ru(1)–O(9)	85.4(2)	O(7)–Ru(2)–C(3)	94.4(5)
Cl(2)–Ru(1)–Cl(3)	172.3(2)	O(7)–Ru(2)–C(8)	88.7(4)
Cl(2)–Ru(1)–C(1)	95.5(6)	C(3)–Ru(2)–C(8)	91.4(8)
Cl(2)–Ru(1)–C(2)	87.6(6)		
Cl(2)–Ru(1)–O(9)	84.2(3)	O(9)–Ru(3)–Cl(1) ⁱ	88.3(3)
Cl(3)–Ru(1)–C(1)	92.2(6)	O(9)–Ru(3)–C(5)	100.2(6)
Cl(3)–Ru(1)–C(2)	93.4(6)	O(9)–Ru(3)–C(7)	87.2(5)
Cl(3)–Ru(1)–O(9)	88.0(3)	O(9)–Ru(3)–C(6)	174.4(7)
C(1)–Ru(1)–C(2)	88.8(6)	O(9)–Ru(3)–O(8)	82.2(4)
C(1)–Ru(1)–O(9)	174.9(4)	Cl(1) ⁱ –Ru(3)–C(5)	87.6(6)
C(2)–Ru(1)–O(9)	96.2(6)	Cl(1) ⁱ –Ru(3)–C(7)	175.2(5)
		Cl(1) ⁱ –Ru(3)–C(6)	92.4(5)
Cl(2)–Ru(2)–O(9)	81.2(3)	Cl(1) ⁱ –Ru(3)–O(8)	93.1(3)
Cl(2)–Ru(2)–C(4)	89.7(5)	C(5)–Ru(3)–C(7)	91.6(7)
Cl(2)–Ru(2)–O(7)	89.3(3)	C(5)–Ru(3)–C(6)	85.4(8)
Cl(2)–Ru(2)–C(3)	99.3(6)	C(5)–Ru(3)–O(8)	177.6(5)
Cl(2)–Ru(2)–C(8)	169.2(6)	C(7)–Ru(3)–C(6)	92.3(7)
O(9)–Ru(2)–C(4)	98.1(6)	C(7)–Ru(3)–O(8)	88.0(5)
O(9)–Ru(2)–O(7)	81.5(4)	C(6)–Ru(3)–O(8)	92.2(6)
O(9)–Ru(2)–C(3)	175.9(6)		
Bond Angles in the Ligands			
Ru(1)–Cl(1)–Ru(3) ⁱ	130.0(2)	Ru(2)–O(7)–C(7)	124.7(8)
Ru(1)–Cl(2)–Ru(2)	88.7(1)	Ru(3)–C(7)–O(7)	117.2(8)
Ru(1)–C(1)–O(1)	176.1(19)	Ru(3)–C(7)–C(9)	126.7(12)
Ru(1)–C(2)–O(2)	178.1(17)	O(7)–C(7)–C(9)	116.1(13)
Ru(2)–C(3)–O(3)	174.8(16)	C(7)–C(9)–C(10)	110.9(12)
Ru(2)–C(4)–O(4)	176.6(17)	Ru(3)–O(8)–C(8)	124.7(10)
Ru(3)–C(5)–O(5)	172.5(15)	Ru(2)–C(8)–O(8)	117.5(11)
Ru(3)–C(6)–O(6)	177.6(13)	Ru(2)–C(8)–C(11)	127.5(12)
Ru(1)–O(9)–Ru(2)	106.0(4)	O(8)–C(8)–C(11)	115.0(14)
Ru(1)–O(9)–Ru(3)	126.1(4)	C(8)–C(11)–C(12)	116.3(16)
Ru(2)–O(9)–Ru(3)	103.0(4)		
Bond angles in the Benzene Molecule			
C(14)–C(13)–C(15)	118.6(31)		
C(13)–C(14)–C(15) ⁱⁱ	119.1(29)		
C(13)–C(15)–C(14) ⁱⁱⁱ	122.3(33)		

^aThe numbers in parentheses are the corresponding standard deviations.

triphenylphosphine or pyridine, with releasing of carbon monoxide and breaking of the bridging structure, the same mononuclear compounds Ru(COC₂-H₅)(CO)L₂Cl·Sv (L = PPh₃, Py; Sv = benzene, methylene chloride, tetrahydrofuran) [2].

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